

# molarity of h2so4

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## Determine the formula of A. Molarity B. Molality C. Mole ... - Socratic

Well, "molarity" is simply the quotient... "molarity" = "moles of solute" / "volume of solution" And "molality" = "moles of solute" / "kilograms of solvent"... For most, dilute, AQUEOUS solutions, "molality" = "molarity". And "mole fraction". is given by the quotient...  $\chi_{\text{the mole fraction}} = \frac{\text{Moles of component}}{\text{Total moles present in the mixture/solution}}$  And "normality" is a bit of an old-fashioned ...

## 255 ml of 0.711 M HCl is mixed with 432 ml of 0.245 M HCl is ... - Socratic

255 ml of 0.711 M HCl is mixed with 432 ml of 0.245 M HCl is mixed with 432 ml of 0.245 M Sr(OH)<sub>2</sub>. Is the solution acid, basic, neutral? if the solution is not neutral, how much of the excess reagent is leftover? What is the Molarity of the excess ? Chemistry

## Question #a14ed - Socratic

Since molarity is defined as the number of moles of solute per liter of solution, increasing the volume while keeping the number of moles of solute constant will result in a decrease in concentration.

## Question #7a784 - Socratic

Expressed in milliliters and rounded to two sig figs, the number of sig figs you have for the molarity of the sulfuric acid solution, the answer will be #V = color (green) ("170 mL")# SIDE NOTE Many textbooks and online sources still use the old definition of STP conditions, at which pressure is #1 atm# and temperature is #0^@"C"#.

## Question #9c19a - Socratic

Every time you're looking for a solution's molarity, you must determine how many moles of solute you get in one liter of solution. That is what molarity essentially tells you - how many moles of solute you'd get if you had exactly one liter of solution.

## Question #5ab95 - Socratic

A solution's molarity is defined as the number of moles of solute, in your case potassium periodate, per liter of solution. So, a 0.03-M solution would contain 0.03 moles of solute for every 1 liter of solution.

## **What is the molarity of a solution prepared by dissolving 3. ... - Socratic**

What is the molarity of a solution prepared by dissolving 3.14 moles of NaOH in water and diluting to a total volume of 4.0 L?

## **What is the molarity of a solution prepared by dissolving 2. ... - Socratic**

What is the molarity of a solution prepared by dissolving 2.350 g KI in enough water to make 175.0 mL of solution ?

## **How does titration affect molarity? + Example - Socratic**

Titration affects molarity by changing the volume of the solution. For example, if you have a 1M solution of NaOH in 1 Liter, then you have 1M Na<sup>+</sup> ions. If you titrate this with 1M HCl, then it will take 1 Liter of HCl to fully titrate this. However, your Na<sup>+</sup> ion concentration is now 1 mole in 2 liters, or 0.5M. So titrations generally increase volume and reduce molarities

## **Question #1565c - Socratic**

The molarity of the silver nitrate solution is 0.394 mM. The balanced chemical equation for this double replacement reaction looks like this  $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$  Notice that you have a 1:1 mole ratio between silver nitrate and potassium chloride, which means that you need 1 mole of silver nitrate for every 1 mole of potassium chloride in order for the reaction to take ...