

# lewis dot structure for ccl4

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## Question #63258 - Socratic

The Lewis structure looks like this: Each carbon atom has 4 unpaired electrons available for bonding. Each of the 3 chlorine atoms has a single unpaired electron available for bonding. The hydrogen has a single electron available for bonding. This results in carbon forming 4 single covalent bonds. The systematic name is trichloromethane.

## What would be the electron dot structure of a molecule of ... - Socratic

In an S<sub>8</sub> ring...? Well, each sulfur atom is sp<sup>3</sup>-"hybridized" to a first approx. And so each sulfur atom has TWO lone pairs of electrons.. In a smaller way we see the same sort of bonding in peroxides...i.e. O-O linkage...

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