

h2co3 lewis dot

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What is the name for the compound H₂CO₃? - Answers

H₂CO₃ is carbonic acid. Carbonic acid is a weak acid that is excreted by the lungs. It also occurs in nature, and leads to formations of stalactites and stalagmites.

Is H₂CO₃ (carbonic acid) soluble or insoluble in water?

Yes, it is soluble. For example, the oceans are constantly getting CO₂ from the atmosphere that gets converted into carbonic acid and thus decreasing pH on a global scale. Or the soft drinks we drink contain carbonic acid in them in dissolved form only due to mixing of CO₂ gas. Also, not all carbonates are insoluble as Ben Norris has commented. Hard and fast 'rules' in ...

What is the full name of the acid H₂CO₃? - Answers

What is the full name of the acid H₂CO₃? - Answers Subjects > Science > Chemistry

How to derive composite acidity constant for H₂CO₃*?

How to derive composite acidity constant for H₂CO₃*? Ask Question Asked 5 years, 8 months ago
Modified 5 years, 8 months ago

What is the Conjugate base of H₂CO₃? - Answers

The conjugate base of H₂CO₃ is HCO₃⁻. When H₂CO₃ donates a proton, it forms the bicarbonate ion (HCO₃⁻), resulting in the conjugate base of the acid.

Which make HCO₃⁻ to show two pH values at two scenarios?

What about the titration of Na₂CO₃ (Same concentration of your considered H₂CO₃) with HCl at the first equivalence point (after added 10ml of HCl). Will the pH same to 8.34 as this? Or If it is not then what is the reason to having different pH values?

What is the pH of H₂CO₃? - Answers

H₂CO₃, also known as carbonic acid, is important in regulating the pH of blood and other bodily fluids. It also plays a role in the transport of carbon dioxide from tissues to the lungs for excretion.

What is conjugate base of H₂CO₃? - Answers

The conjugate base of H₂CO₃ is HCO₃⁻. It is formed when H₂CO₃ donates a proton (H⁺) in a reaction. You mean, HCO₃⁻ = bicarbonate H₂CO₃ = carbonic acid and the conjugate of the above base.

Is H_2S a strong electrolyte? - Answers

No, carbonic acid (H_2CO_3) is a weak acid, as it only partially dissociates in water solution. Thus, it is not a strong electrolyte.

The instability of hydrated carbon dioxide or "carbonic acid"

Note that waterfree H_2CO_3 is reportedly kinetically very stable, as solid or gas, subliming near -55°C . A molecule of H_2CO_3 has reportedly half-life 180 000 years. But there is extremely strong autocatalytic effect of water. // At ambient temperatures, pure carbonic acid is a stable gas. [5]
There are two main methods to produce anhydrous carbonic acid: reaction of hydrogen chloride and ...