

# **h<sub>2</sub>co<sub>3</sub> lewis dot**

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## **What is the name for the compound H<sub>2</sub>CO<sub>3</sub>? - Answers**

H<sub>2</sub>CO<sub>3</sub> is carbonic acid. Carbonic acid is a weak acid that is excreted by the lungs. It also occurs in nature, and leads to formations of stalactites and stalagmites.

## **Is H<sub>2</sub>CO<sub>3</sub> (carbonic acid) soluble or insoluble in water?**

Yes, it is soluble. For example, the oceans are constantly getting CO<sub>2</sub> from the atmosphere that gets converted into carbonic acid and thus decreasing pH on a global scale. Or the soft drinks we drink contain carbonic acid in them in dissolved form only due to mixing of CO<sub>2</sub> gas. Also, not all carbonates are insoluble as Ben Norris has commented. Hard and fast 'rules' in ...

## **What is the full name of the acid H<sub>2</sub>Co<sub>3</sub>? - Answers**

What is the full name of the acid H<sub>2</sub>Co<sub>3</sub>? - Answers Subjects > Science > Chemistry

## **How to derive composite acidity constant for H<sub>2</sub>CO<sub>3</sub>\*?**

How to derive composite acidity constant for H<sub>2</sub>CO<sub>3</sub>\*? Ask Question Asked 5 years, 8 months ago Modified 5 years, 8 months ago

## **What is the Conjugate base of H<sub>2</sub>co<sub>3</sub>? - Answers**

The conjugate base of H<sub>2</sub>CO<sub>3</sub> is HCO<sub>3</sub><sup>-</sup>. When H<sub>2</sub>CO<sub>3</sub> donates a proton, it forms the bicarbonate ion (HCO<sub>3</sub><sup>-</sup>), resulting in the conjugate base of the acid.

## **Which make HCO<sub>3</sub><sup>-</sup> to show two pH values at two scenarios?**

What about the titration of Na<sub>2</sub>CO<sub>3</sub> (Same concentration of your considered H<sub>2</sub>CO<sub>3</sub>) with HCl at the first equivalence point (after added 10ml of HCl). Will the pH same to 8.34 as this? Or If it is not then what is the reason to having different pH values?

## **What is the pH of H<sub>2</sub>CO<sub>3</sub>? - Answers**

H<sub>2</sub>CO<sub>3</sub>, also known as carbonic acid, is important in regulating the pH of blood and other bodily fluids. It also plays a role in the transport of carbon dioxide from tissues to the lungs for excretion.

## **What is conjugate base of H<sub>2</sub>CO<sub>3</sub>? - Answers**

The conjugate base of H<sub>2</sub>CO<sub>3</sub> is HCO<sub>3</sub><sup>-</sup>. It is formed when H<sub>2</sub>CO<sub>3</sub> donates a proton (H<sup>+</sup>) in a reaction. You mean, HCO<sub>3</sub><sup>-</sup> = bicarbonate and H<sub>2</sub>CO<sub>3</sub> = carbonic acid and the conjugate of the above base.

## **Is h<sub>2</sub>s a strong electrolyte? - Answers**

No, carbonic acid (H<sub>2</sub>CO<sub>3</sub>) is a weak acid, as it only partially dissociates in water solution. Thus, it is not a strong electrolyte.

## **The instability of hydrated carbon dioxide or "carbonic acid"**

Note that waterfree H<sub>2</sub>CO<sub>3</sub> is reportedly kinetically very stable, as solid or gas, subliming near -55 deg C. A molecule of H<sub>2</sub>CO<sub>3</sub> has reportedly halflife 180 000 years. But there is extremely strong autocatalytic effect of water. // At ambient temperatures, pure carbonic acid is a stable gas. [5]

There are two main methods to produce anhydrous carbonic acid: reaction of hydrogen chloride and

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