

hco3 strong or weak

AI generated article from Bing

HCO3 = CaCO3? - Homebrew Talk

My water report lists "bicarbonate as CaCO3" but BeerSmith is asking for bicarbonate as HCO3. Can one derive the former from the latter? Does any of this really matter since apparently I should be using EZ Water calculator? Also, is the convention to use averages when determining the constituents of one's water?

Adding baking soda to the kettle to achieve HCO3 target?

Since there really isn't a beer style HCO3 ppm target, the purpose of baking soda, or any source of HCO3, is solely to raise a low mash PH to the ideal mash PH - typically only needed for dark styles. The amount of baking soda added to the mash should be based on what is required to hit the target mash PH.

Total Alkalinity, CaCO3 as a function of Bicarbonate, HCO3

My most recent Lab Wards report shows my Total Alkalinity, CaCO3 as 89, my Bicarbonate, HCO3 as 97 and my pH as 8.1. I had read that for pH under 8.3 Total Alkalinity, CaCO3 = 50/61 * Bicarbonate, HCO3. That does not hold true for my numbers. My report also lists Carbonate, CO3 as 6. Does...

Reducing HCO3 - Homebrew Talk

I'm from Galveston Texas. I have two broad questions, but bear with some background information first: 2016 Water report: Ca 2+ 37 Mg 2+ 7 HCO3 - 127 SO4 2- 38 Na + 37 Cl - 51 pH 8.1 (their measurement) pH 7 (my presumably much less accurate measurement at the faucet) I only...

alkalinity as CaCO3 - Homebrew Talk

This assumes that HCO3- has one alkalinity equivalent and that CaCO3 has 2 alkalinity equivalents. As a result you multiply [HCO3-] with 50 and divide by 61. The result is 0.5ppm alkalinity as CaCO3. But there is a problem. The actual alkalinity contributed by CaCO3 was not one equivalent but 2.

Bicarbonate from Total Alkalinity - Homebrew Talk

To quote him from an email: "HCO3 (bicarbonate) is a fractional part of CaCO3 (units of the total alkalinity), by means of the molecular weights 100.086 to 61.016 (61.016/100.086 = 0.6096). This concerns mass units only. This says HCO3 is 60.96% of the mass of CaCO3. The titration test for these is the same, using Sulfuric Acid.

Total Hardness vs Alkalinity - Homebrew Talk

bicarbonate == HCO_3 , so take a wild guess ;-) Alkalinity "as CaCO_3 " is a fictional unit and means how much alkalinity you'd get if you added that much CaCO_3 . If you add for example potassium carbonate (K_2CO_3), you'd still get "alkalinity as CaCO_3 ", but you wouldn't have added any CaCO_3 .

Very high bicarbonate in water, what to do? - Homebrew Talk

Hi all, I just got a water analysis from the town council and entered the numbers into my software. It turns out that the HCO_3 level is 397ppm. As I understand, that reduces the drop of PH while mashing. what else I've learned is that a can 1. Dilute the water with distilled water (but won't...

brewing with very hard water - reducing bicarbonate

HCO_3^- has a 'rounded' molecular weight of 61 and a valence of -1 CaCO_3 has a 'rounded' molecular weight of 100, and its cation and anion components have valences of +2 and -2 respectively.

Calculating HCO_3^- level from other parameters. - Homebrew Talk

Ok, so it's been a little while since high school chemistry so I need a little help. I got a water profile from the town which had all of the necessary info to plug into Beersmith 2 except the HCO_3^- . What the report does include is two parameters which I think should allow me to calculate...