

# ccl4 lewis dot structure

*AI generated article from Bing*

---

**Answered: Methane (CH<sub>4</sub>) is a gas, but carbon tetrachloride ... - bartleby**

CH<sub>4</sub> exhibits hydrogen bonding, but CCl<sub>4</sub> does not, therefore the bp of CCl<sub>4</sub> is higher. O CCl<sub>4</sub> is a polar molecule and CH<sub>4</sub> is nonpolar, therefore CCl<sub>4</sub> has a lower bp. Methane (CH<sub>4</sub>) is a gas, but carbon tetrachloride (CCl<sub>4</sub>) is a liquid at room temperature and standard pressure conditions. Which one of the following is the best explanation for this?

**Calculate  $\Delta H_{\text{rxn}}$  for the following reaction: CH<sub>4</sub> (g)+4Cl<sub>2</sub> (g)→CCl<sub>4</sub> (g) ...**

Calculate  $\Delta H_{\text{rxn}}$  for the following reaction: CH<sub>4</sub> (g)+4Cl<sub>2</sub> (g)→CCl<sub>4</sub> (g)+4HCl (g) given these reactions and their  $\Delta H$  values: C (s)+2H<sub>2</sub> (g)→CH<sub>4</sub> (g),  $\Delta H = -74.6 \text{ kJ}$  C (s)+2Cl<sub>2</sub> (g)→CCl<sub>4</sub> (g),  $\Delta H = -95.7 \text{ kJ}$  H<sub>2</sub> (g)+Cl<sub>2</sub> (g)→2HCl (g),  $\Delta H = -184.6 \text{ kJ}$  Express the enthalpy in kilojoules to one decimal place.

**Answered: CH<sub>4</sub> (g) + 4Cl<sub>2</sub> (g) → CCl<sub>4</sub> (g) + 4HCl (g),  $\Delta H$  ... - bartleby**

Solution for CH<sub>4</sub> (g) + 4Cl<sub>2</sub> (g) → CCl<sub>4</sub> (g) + 4HCl (g),  $\Delta H = -434 \text{ kJ}$  Based on the above reaction, what energy change occurs when 2.2 moles of methane (CH<sub>4</sub>) reacts?

**Answered: Methane gas (CH<sub>4</sub>) reacts with chlorine gas (Cl<sub>2</sub>) ... - bartleby**

Solution for Methane gas (CH<sub>4</sub>) reacts with chlorine gas (Cl<sub>2</sub>) to produce liquid carbon tetrachloride (CCl<sub>4</sub>) and hydrogen chloride gas (HCl). The balanced...

**Calculate  $\Delta H_{\text{rxn}}$  for the following reaction: CH<sub>4</sub> (g)+4Cl<sub>2</sub> (g)→CCl<sub>4</sub> (g) ...**

Solution for Calculate  $\Delta H_{\text{rxn}}$  for the following reaction: CH<sub>4</sub> (g)+4Cl<sub>2</sub> (g)→CCl<sub>4</sub> (g)+4HCl (g) given these reactions and their  $\Delta H$  values: C (s)+2H<sub>2</sub> (g) →CH<sub>4</sub>,...

**Identify the best reagents to convert 1-hexyne into 2-bromo ... - bartleby**

Solution for Identify the best reagents to convert 1-hexyne into 2-bromo-1-hexene. O xs Br<sub>2</sub>, CCl<sub>4</sub> 1 equiv HBr, ROOR xs HBr O 1 equiv HBr O 1 equiv. Br<sub>2</sub>, CCl<sub>4</sub>

## Identify the best reagents to convert 1-hexyne into 2,2 ... - bartleby

Solution for Identify the best reagents to convert 1-hexyne into 2,2-dibromohexane. O 1 equiv. Br<sub>2</sub>, CCl<sub>4</sub> O 1 equiv HBr, ROOR O 1 equiv HBr Oxs Br<sub>2</sub>, CCl<sub>4</sub> Oxs HBr...

## Answered: Draw the lewis structure of carbon tetrachloride (CCl4 ...

Solution for Draw the lewis structure of carbon tetrachloride (CCl<sub>4</sub>). include all valance electrons. name the electron geometry and molecular geometry for this...

## Answered: Carbon tetrachloride (CCl4) and benzene (C6H6 ... - bartleby

Carbon tetrachloride (CCl<sub>4</sub>) and benzene (C<sub>6</sub>H<sub>6</sub>) form ideal solutions. Consider an equimolar solution of CCl<sub>4</sub> and C<sub>6</sub>H<sub>6</sub> at 25°C. The vapor above the solution is collected and condensed. Using the following data, determine the composition in mole fraction of the condensed vapor.

## determine the milligrams of CCl4 produced with 325 milligrams of CS2 ...

Solution for determine the milligrams of CCl<sub>4</sub> produced with 325 milligrams of CS<sub>2</sub> reacts with CL<sub>2</sub> according to the following reaction: CS<sub>2</sub> (g) +...