bh3 lewis structure polar or nonpolar

AI generated article from Bing

Order of lewis acidity for BBr3, BCl3, BMe3, BH3, BPh3?

BBr3 is more acidic than BCl3 due to weaker pi back bonding in BBr3 as compared to BCl3.BCl3 is stronger acid than BH3 due to electron withdrawing effect of Cl make B becomes more electron deficient despite of pi back bonding in BCl3. Due to steric reason BMe3 is stronger acid than BPh3.

organic chemistry - Why does BH3/THF reduce the carbonyl group to a ...

Why does BH3/THF reduce the carbonyl group to a methylene rather than an alcohol? Ask Question Asked 3 years, 11 months ago Modified 3 years, 11 months ago

Why do BH3 and NaBH4 have different selectivities?

I've always known borane ($\\ce{BH3}$), as a reagent for alkene hydroboration. Recently in class we talked about its use in the reduction of carboxylic acids to alcohols as well. Now obviously $\\ce...

organic chemistry - What is the balanced equation of BH3/THF with an ...

The reduction of secondary amides with BH3 requires 3 equivalents, mechanism shown below image sourced from here. The borane becomes HB=O which may transform further depending on work up.

Reduction of acids with borane - Chemistry Stack Exchange

What is the mechanism of the reduction of organic acids with borane? Borane reductions don't work well with electron deficient carbonyl groups. Why is that so, and how isn't the carboxyl group an

hybridization - What makes banana bonds possible in diborane ...

Diborane has the interesting property of having two 3-centered bonds that are each held together by only 2 electrons (see the diagram below, from Wikipedia). These are known as "banana bonds." ...

How does a hydroboration mechanism actually work?

In BH3, Boron is less electronegative than H (i.e. hydrogen) , hence in the syn addition I have shown, hydrogen ends up being attached to a more substituted carbon atom.

Why BH3 its classified as a "Molecular Hydride" and AlH3 its a ...

Why BH3 its classified as a "Molecular Hydride" and AlH3 its a "Intermediary Hydride" Ask Question Asked 8 years, 7 months ago Modified 8 years, 7 months ago

Why can't sodium borohydride reduce carboxylic acid and ester groups ...

Sodium borohydride (NaBHX4 N a B H X 4) is a reducing agent used in various organic reaction to reduce aldehydes, ketones and acyl chlorides. Why can't sodium borohydride reduce carboxylic acid groups and esters? Why can it reduce acyl chlorides but not esters, both which are derivatives of the carboxylic acid group? I read somewhere that it had to do with the oxidizing property of sodium ...

Dipole moment of symmetrical molecules - Chemistry Stack Exchange

While studying dipole moment in organic chemistry, I got stuck at a particular statement written in the book : Symmetrical molecules without lone pairs of electrons will have $\\mu = 0.$ The main